

Introduction à la cinétique électrochimique

Figure 1 : Couple rapide

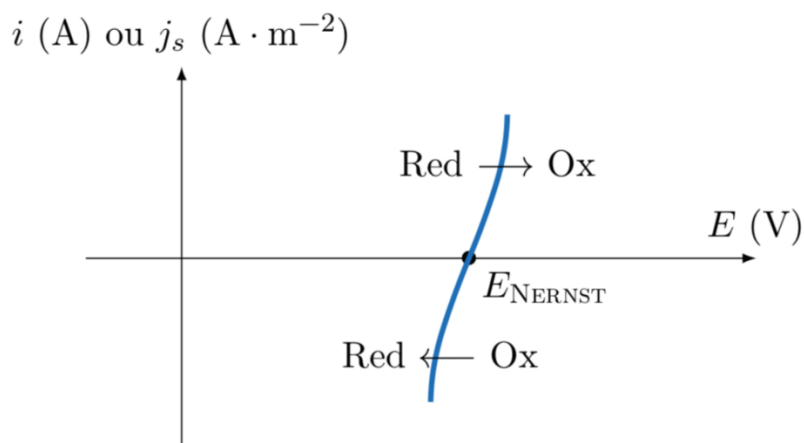


Figure 2 : Couple lent

i (A) ou j_s ($A \cdot m^{-2}$)

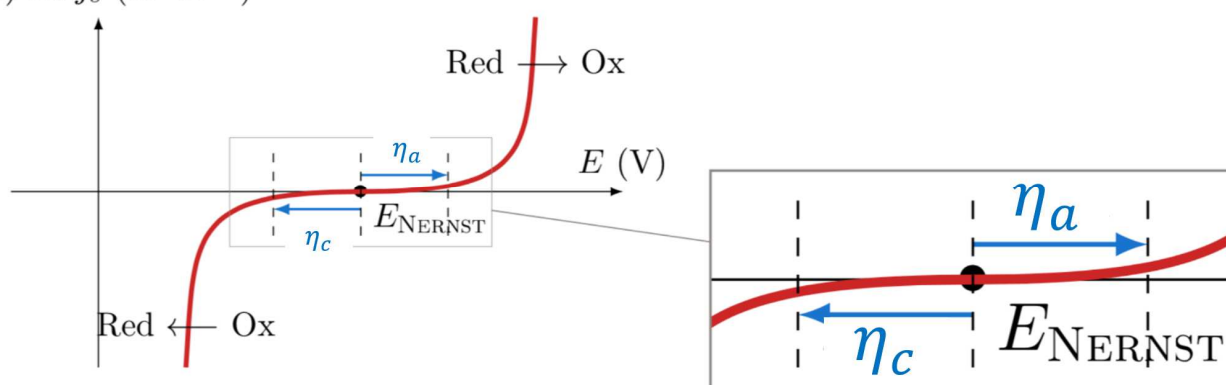


Figure 3 : Nature de l'électrode

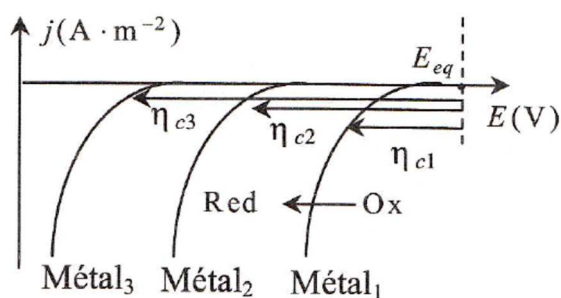


Figure 4 : palier de diffusion

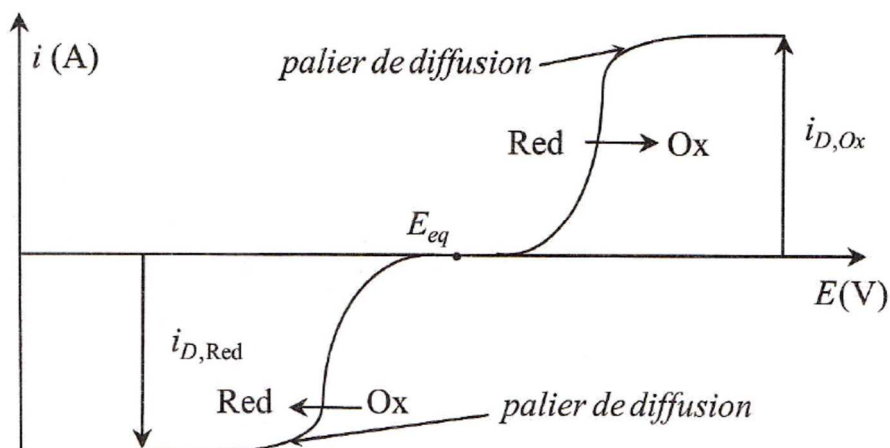


Figure 5 : vagues successives

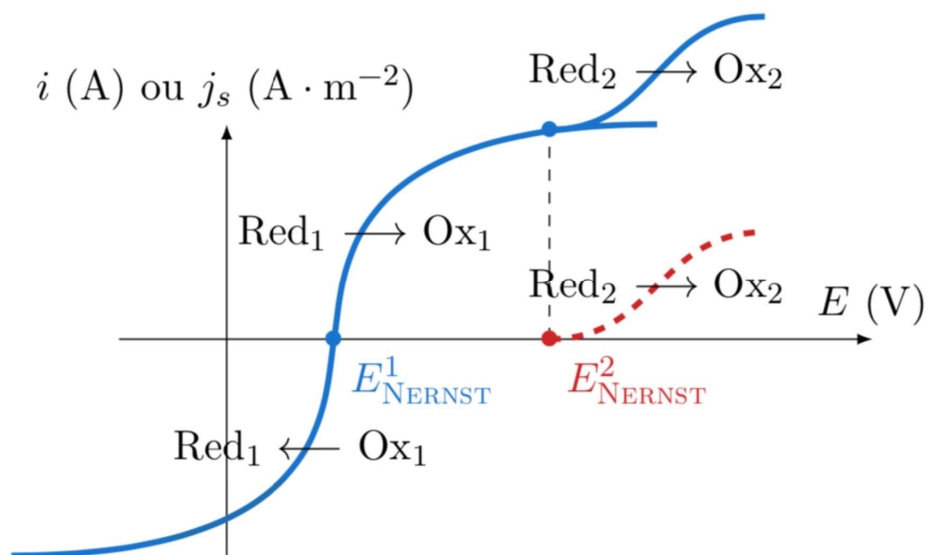


Figure 6 : mur du solvant

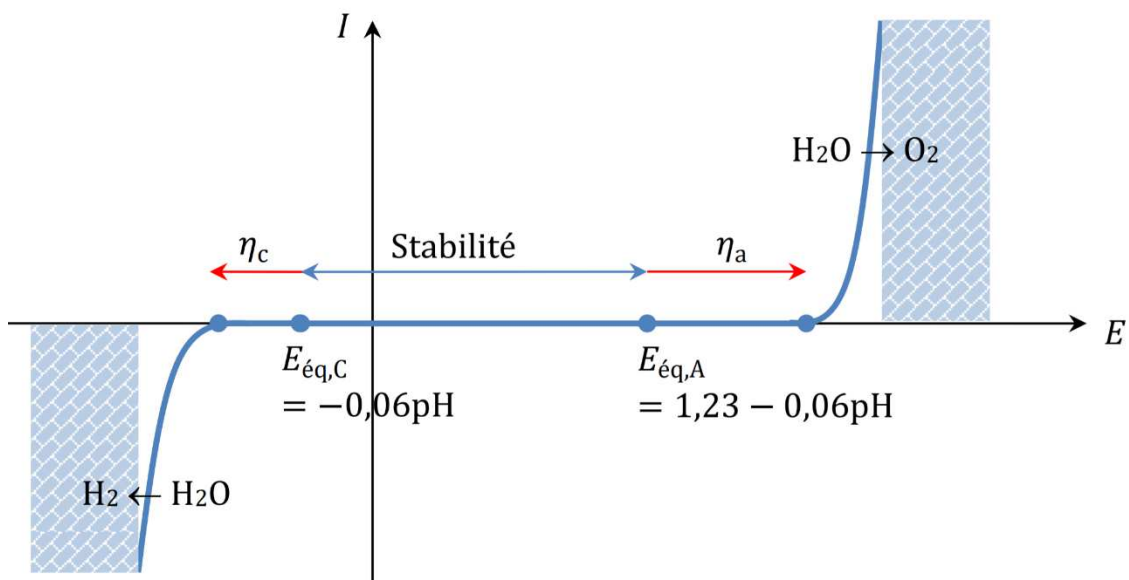


Figure 7 : cas général couple Ox/Red

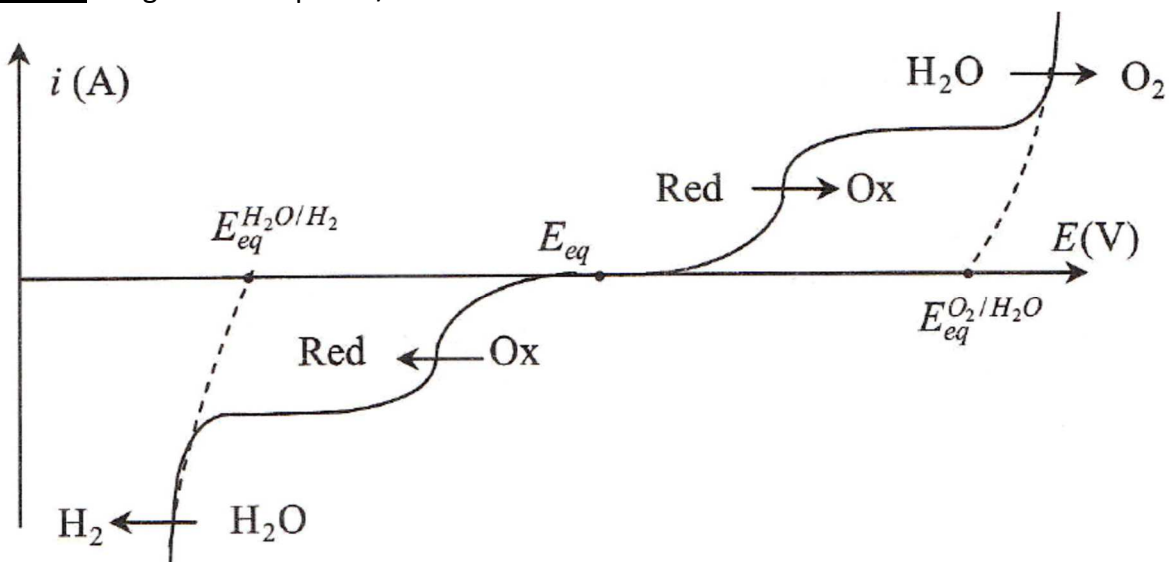
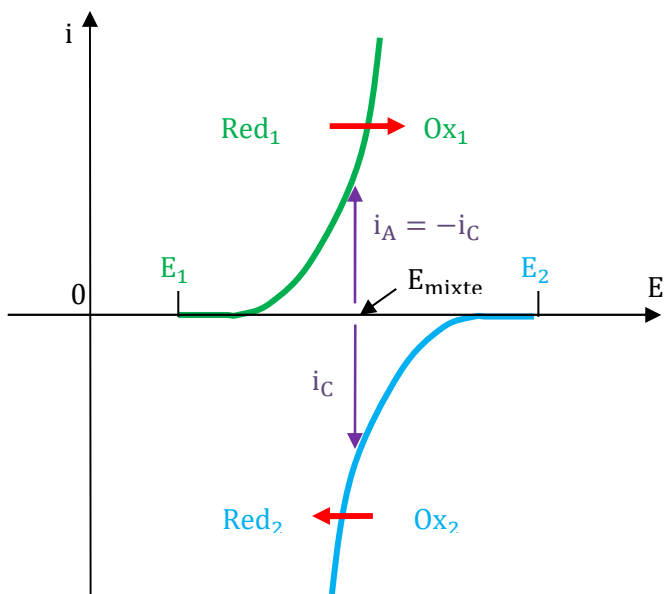
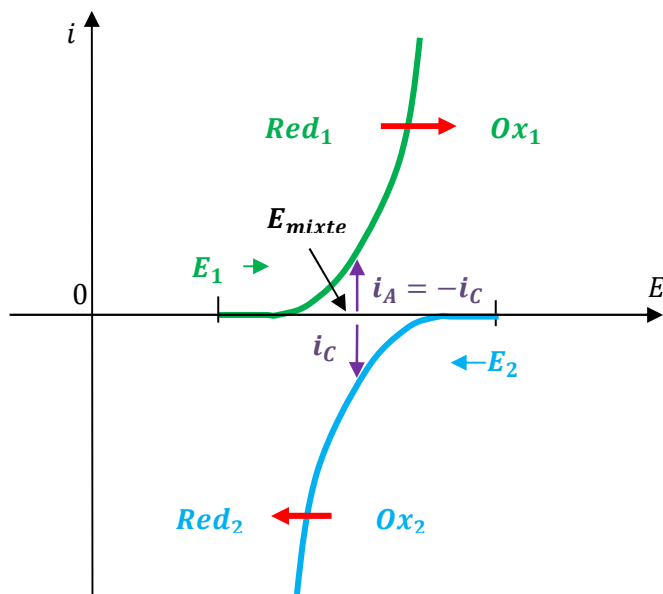


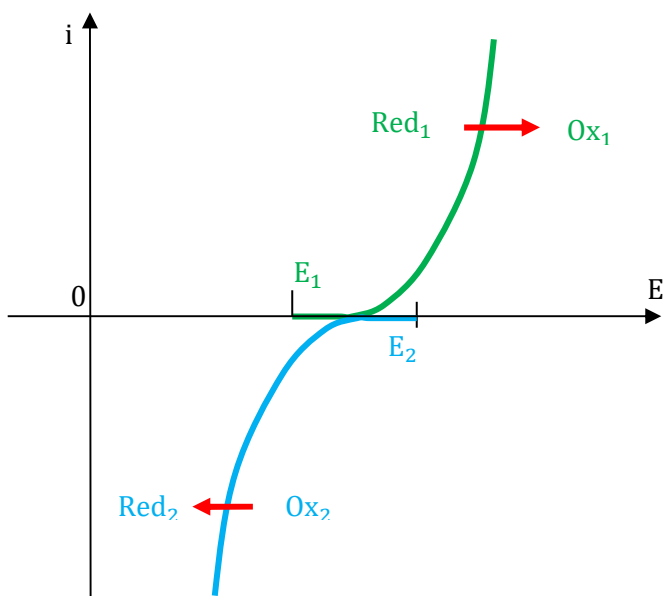
Figure 8 : évolution d'une réaction thermodynamiquement possible au cours du temps



Réaction cinétiquement rapide ($t=0$)



Réaction cinétiquement lente (à $t_1 > 0$)



Réaction cinétiquement bloquée (à $t_2 > t_1$)